IDEAL GAS LAW

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**Abstract:**

**Purpose:**

**Relevant Findings:**

Flask a + Manginess dioxide: 84.6319g Manginess dioxide: 0.02g

Flask a + Potassium Chloride: 84.8427g

Flask b: 82.830g

Flask b with water: 134.123g

Temperature in flask b: 21.63

Pressure flask b: 0.9889 atm

End flask a: 84.7756g

**Part 1**

To Determine the mass of oxygen produced during the reaction, first we need to balance the equation: